

**Dissolved Lead – values in pmol/kg  
Consensus values ( $\pm$  1 std. dev.) for North Atlantic GEOTRACES  
Reference Samples as of May 2013**

**GEOTRACES GS = 28.6  $\pm$  1.0 pmol/kg**

**GEOTRACES GD = 42.7  $\pm$  1.5 pmol/kg**

These are considered to be the consensus values for the GEOTRACES reference samples as of May 2013.

**Labs participating in the analysis of the North Atlantic GEOTRACES  
reference samples to determine a consensus value for dissolved Pb:**

**Yoshiki Sohrin (U Kyoto, Japan):**

Off line concentration using an EDTriA-type chelating resin with subsequent analyses by ICP-MS using the method of Sohrin et al. (2008).

**Michael Ellwood (Australian National U, Australia):**

Lead was concentrated by solvent extraction (Bruland et al., 1979) and analyzed by ICPMS. 100 g seawater samples were buffered to a pH of 4.5 with purified ammonium acetate buffer. Purified ammonium pyrrolidinedithiocarbamate (PDC) and sodium diethyldithiocarbamate (DDC) were added to the samples which were then extracted twice by shaking following the addition of purified chloroform. The two chloroform extracts obtained were combined, acidified with nitric acid, shaken for 1 min and then diluted with purified water. Trace metal concentrations were determined by ICP-MS.

**Ana Aguilar-Islas/Jingfeng Wu (UAF, U.S.):**

Mg(OH)<sub>2</sub> coprecipitation and analysis by isotope dilution ICP-MS (Wu and Boyle, 1997). Lead was analyzed by isotope dilution using the ratio between the natural abundance of <sup>208</sup>Pb and an added <sup>204</sup>Pb spike. Interference from <sup>204</sup>Hg was corrected by monitoring <sup>202</sup>Hg counts. 1.6ml of acidified sample and spike were allowed to equilibrate for several minutes. A single coprecipitation step was carried out followed by dilution of the precipitate with 4% HNO<sub>3</sub>. Blanks were done using 50 ul of low Pb seawater instead of 1.6ml.

**Ed Boyle (MIT, U.S.):**

400-bead NTA-type resin with isotope dilution ICP-MS (Lee et al. 2011). The 1.3 ml samples in 1.5 cc microcentrifuge tubes were spiked with <sup>204</sup>Pb, the solution pH was raised to 6 using ammonium acetate. The beads were added, and left to equilibrate overnight on a shaker table. The samples were centrifuged and supernatant siphoned off. The beads were washed/centrifuged/siphoned three times with high purity distilled water to eliminate salt. 150  $\mu$ l of 0.1N HNO<sub>3</sub> was added and allowed at least one day to release the Pb into the acid. The Pb 206/204 ratio was then determined by quadrupole ICP-MS using a low-flow micromist nebulizer.

**Jingfeng Wu (Univ. of Miami, U.S.):**

Mg(OH)<sub>2</sub> coprecipitation and analysis by isotope dilution ICP-MS (Wu and Boyle, 1997).

**Peter Croot/Peter Streu (IMF/GEOMAR, Germany):**

Samples were analyzed according to the method described in Kremling and Streu (2001). For the analysis of Cd, Co, Cu, Fe, Ni, Pb and Zn, 300–500 g portions of the samples were subjected to a dithiocarbamate–freon extraction modified from the procedure by Danielsson et al. (1978) implying maximum concentration factors of 500. The final extracts with the metals were measured by electrothermal atomic absorption spectrometry with Zeeman background correction (ETAAS; Perkin-Elmer Model 4100 ZL).

**Christa Pohl (Warnemunde, Germany):**

Samples were analyzed according to the method described in Kremling and Streu (2001). The final extracts with the metals were measured by electrothermal atomic absorption spectrometry.

**Dondra Biller/Ken Bruland (UCSC, U.S.):**

Off-line concentration using an EDTri-A-type chelating resin with subsequent analyses by ICP-MS (Biller and Bruland, submitted) based upon the method of Sohrin et al. (2008). The method entails an eight column manifold enabling eight separate ~ 40 mL samples to be processed simultaneously (Biller and Bruland, 2012).

**Angie Milne/Bill Landing (FSU, U.S.):**

Off-line extraction using IDA Toyopearl AF-Chelate resin followed by analysis using isotope dilution ICP-MS (Milne et al. 2010). Prior to extraction the samples (12 mL) were UV oxidized and buffered to pH ~6.2.

**Geoff Smith/Ken Bruland (UCSC, U.S.):**

On-line flow injection analysis of 4 ml of sea water using an EDTri-A-type chelating resin (Sohrin et al., 2008) at pH 6 utilizing purified ammonium acetate buffer and eluting analytes with 1.5M HNO<sub>3</sub> followed by detection with ICPMS.

**Eric Achterberg (Plymouth Univ., UK):**

Off-line extraction using IDA Toyopearl AF-Chelate resin followed by analysis using isotope dilution ICP-MS (Milne et al. 2010).

**Celine Gallon, Cheryl Zurbrick and Russ Flegal (UCSC, U.S.):**

On-line extraction using IDA Toyopearl AF-Chelate resin followed by analysis using isotope dilution ICP-MS (Ndung'u et al. 2003).

**Christian Schlosser and Eric Achterberg (Plymouth, UK)**

Off-line extraction using a WAKO chelating resin (Kagaya, 2009) followed by analysis on an Element XR ICP-MS. Samples were UV digested for 3 hours.

**Rob Middag and Ken Bruland (UCSC, US)**

Off-line extraction with Nobias PA-1 chelating resin and analysis on an Element XR ICP-MS (Middag et al., submitted).

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